

Roll No. _____

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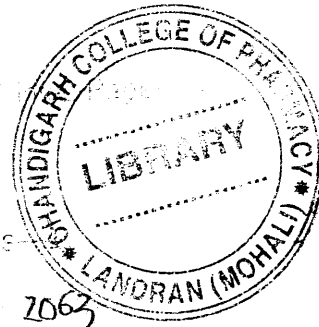
Q. No. of Questions _____

Total No. of _____

1.1.1

Pharmaceutical Analysis

(B. Pharmacy 1st Semester) 2063



Time's Hour _____

Mark's _____

Note :- This paper consists of Three sections, section A to be compulsory, answer any Two questions from Section B and any Three questions from Section C

Section-A

- (a) How do you define an assay?
- (b) Define the terms trace component, major and minor constituent
- (c) Calculate and express the result in correct number of significant figures:
 $120.30 + 390.69 + 0.02308$
 $950.0 \times 0.005 \times 1.498$

1.1.1

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- (d) Give the conjugate bases for the following.
HCl, Water.
- (e) Explain the concept of activity and activity coefficient.
- (f) What is K_a and what is its relation with acid strength?
- (g) What is pH of a solution having hydronium ion concentration of $1.5 \times 10^{-4} \text{ mol/l}$?
- (h) Calculate the pH of 0.05 M sodium benzoate solution. pK_a for benzoic acid is 4.2.
- (i) What is colour change interval of an indicator?
- (j) Give the full procedure for the determination of the end-point of the titration of a weak acid with a strong base.
- (k) Write the function of potassium iodide in iodine solution?

1.1.1

Answer

(3)

- (m) Arrange the following in the order of their oxidizing power :
 KMnO_4 , $\text{K}_2\text{Cr}_2\text{O}_7$, Iodine and ceric sulphate.
- (n) Calculate the solubility of silver chloride in 0.01 M NaCl solution. K_s for AgCl is $1.1 \times 10^{-10} \text{ mol}^2/\text{l}^2$.
- (o) What is the use of digestion of precipitates in gravimetry ?

Section-B Marks : 5 Each

2. 20 determinations were made for weight of ascorbic acid in tablet samples. Mean result was 495.2 mg. Calculate whether result is significant in the assay if true value is 488.5 mg and standard deviation is 1.8. Given tabulated t -values are 1.728, 1.729 and 2.539 for probability levels of 0.05 and 0.01 respectively.
3. Describe the types of neutralization curves obtained in the titration of a weak acid with a strong base and a strong acid with a weak base. Which indicators are used in this titration ?

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7. Give pharmaceutical applications of Mohr's method.
8. Solubility of silver chloride is 1.2×10^{-5} g/l. Calculate its solubility product.
9. Explain the thermogravimetric curve of copper sulfate highlighting its salient points.

Section--C (Marks: 10 each)

1. Write down the role of solubility product and the method of determining the end point in Mohr's method of chloride estimation? How does this method differ from Volhard's method?
2. Discuss various types of errors in pharmaceutical analysis.
3. Weigh ten tablets in a weight variation test and got the following data. Calculate mean, standard deviation and coefficient of variation for the data:
200.0, 198.5, 205.5, 195.0, 210.2, 190.5, 192.0, 215.0, 203.5, 202.7.
4. Derive the expression for hydrolysis constant and pH of a solution of a salt of weak acid and strong base.

1.1.1

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