)ll No.

otal No. of Questions: 10]

[Total No. of Pages: 02

J-722 [5367]

[2126]

B.Pharmacy (Semester - 1st)

PHARMACEUTICAL ANALYSIS - I (PHM - 1.1.1)

Time: 03 Hours

Maximum Marks: 80

Instruction to Candidates:

- 1) Section A is **compulsory.**
- 2) Attempt any **Four** questions from Section B.
- 3) Attempt any **Three** questions from Section C.

Section - A

Q1)

 $(15 \times 2 = 30)$

- a) Explain the term 'Accuracy'?
- b) Explain the hydrolysis of salts of strong acid and weak base?
- c) Define 'Molarity' and Explain how do you prepare 1 molar NaOH solution?
- d) Differentiate between primary and secondary standard solutions?
- e) Explain Bronsted acid-base theory?
- f) What is buffer solution? Explain briefly the importance of buffer solutions in pharmacy?
- g) Define common ion effect and give its practical importance?
- h) Explain the concepts of PPM and PPB?
- i) Write notes on universal indicators?
- j) Explain briefly the preparation and standardization of N/10 KMnO₄.
- k) Differentiate between internal and external redox indicators?
- l) Write notes on digestion and peptisation?
- m) Discuss the organic precipitants in gravimetry?

- n) Explain how do you determine calcium as calcium oxalate by gravimetry
- o) Define standard deviation and give its formula?

Section - B

 $(4 \times 5 = 20)$

- Q2) Explain systematic and random errors with suitable examples.
- Q3) Explain briefly the concept of post precipitation?
- **Q4**) Write notes on buffer action and give the significance of Henderson and Hesselbach equation?
- Q5) Discuss the theory of redox titrations?
- Q6) Discuss the classification of chemical reactions involved in volumetric analysis?

Section - C

 $(3 \times 10 = 30)$

- Q7) What are neutralization curves? Explain the titration curves of strong acid and strong base?
- Q8) Discuss the basic steps involved in gravimetry?
- Q9) Explain how the end point is detected in redox titrations?
- Q10) Discuss various sampling techniques used in pharmacy?

