

Roll No.

Total No. of Questions : 10]

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B. Pharmacy (Sem. - 2nd)
PHARMACEUTICAL CHEMISTRY - II
(Physical Chemistry)
SUBJECT CODE : PHM-1.2.3.

Paper ID : [D0109]

[Note : Please fill subject code and paper ID on OMR]

Time : 03 Hours

Maximum Marks : 80

Instruction to Candidates:

- 1) Section - A is **Compulsory**.
- 2) Attempt any **Four** questions from Section - B.
- 3) Attempt any **Three** questions from Section - C.

Section - A

Q1)

(15 x 2 = 30)

- a) Temperature coefficient.
- b) Ideal solution.
- c) Degree of freedom.
- d) Ionic mobility.
- e) Charle's law.
- f) Critical temperature.
- g) Fluidity.
- h) Binary liquid solution.
- i) System.
- j) Grothus Draper law.
- k) Adsorption.
- l) Order of reaction.
- m) Adiabatic process.
- n) Phosphorescence.
- o) Surface tension.

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P.T.O.

Section - B

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- Q2) Derive Beer Lambert law.
- Q3) Define dipole moment. How it is useful for elucidating the molecular structure?
- Q4) What are the postulates of quantum mechanics?
- Q5) Describe the principle of fractional distillation.
- Q6) Describe zeroth law of thermodynamics.

Section - C

- (3 x 10 = 30)
- Q7) What are colligative properties? Describe the colligative properties of dilute solution.
- Q8) Describe Debye-Huckel theory of strong electrolytes.
- Q9) Describe the various methods used to determine the order of a reaction.
- Q10) Describe langmuir theory of adsorption.



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