D.U. M.Sc. CHEMISTRY ENTRANCE - 2017 (SET-A)

1. The oxidation states of oxygen in O_2F_2 is
    (a) +1   (b) +2   (c) +4   (d) –2

2. Which among the following is a powerful oxidizer?
    (a) F_2 (b) Cl_2 (c) Br_3 (d) I_2

3. The most suitable in water among the following is
    (a) $CsClO_4$ (b) $KClO_4$ (c) $NaClO_4$ (d) $LiClO_4$

4. Dental filling is done by
    (a) Portland cement   (b) Sorrel cement
    (c) Calcia   (d) Double salt

5. The polymeric nature of boric acid is due to
    (a) its acidic nature   (b) its gemetry
    (c) presence of hydrogen bonds   (d) monobasic nature

6. Alkali metals form highly stable complexes with
    (a) Cryptand-222   (b) Diethyl ether   (c) Cyclopentadiene   (d) Butadiene

7. Of the following, which is a low expands one?
    (a) Sodaline   (b) Quartz   (c) Vycor   (d) Borosilicate

8. Which of the following has centre of inversion?
    (A) CO_2 (B) C_2H_2 (C) BF_3 (D) SO_4^{3-}
    (a) A and B only   (b) A and C only   (c) A and D only   (d) B and C only

9. "Yellow when hot and white and cold is one of the characteristics observed for ZnO. It is due to
    (a) Distortions   (b) Cation defects   (c) Anion defects   (d) Deformations

10. Which is a superconductor?
    (a) $Bi_2CaSr_2Cu_2O_8$ (b) $Bi_2CaSr_2Co_2O_8$
    (c) $Bi_2CdSr_2Cu_2O_8$ (d) $Bi_2CaSn_2Cu_2O_8$

11. Electrophoresis refers to
    (a) Separation   (b) Identification   (c) Digestion   (d) Amalgamation

12. Find the amphoteric one in the following:
    (a) CO_2 (b) Mn_2O_2 (c) MgO (d) Sb_3O_3

13. The reaction of O_3 with KI produces:
   (a) HI   (b) HIO_8 (c) I_8 (d) I_3^+

14. Which one of the following cannot show linkage isomerism?

    (a) NO_3^-　　　　(b) SCN　　　　(c) CN^-　　　　(d) NH_3

15. A metal X on heating in nitrogen gas gives Y. Y on treatment with H_2O gives a colourless gas which when passed through $CuSO_4$ solution gives blue colour. Y is

    (a) $Mg(NO_3)_2$　　　　(b) Mg_2N_3　　　　(c) $MgCl_2$　　　　(d) MgO

16. Hydrogen as fuel is stored as

    (a) gas　　　　(b) semi liquid　　　　(c) liquid　　　　(d) solid

17. Which of the following order is correct for scattering of X-rays?

    (a) F < O < Cl < Na < Tl　　　　(b) F < Cl < Na < O < Tl

    (c) O < Cl < F < Na < Tl　　　　(d) O < F < Na < Cl < Tl

18. The p-electronic configuration is equivalent to the term:

    (a) ^3P　　　　(b) ^2P　　　　(c) ^3F　　　　(d) ^4P

19. The anticancer drug cisplatin is

    (a) $Pt(NH_3)_2Cl_2$　　　　(b) $Pt(H_2O)_2Cl_2$　　　　(c) $Pt(NH_3)_3Cl$　　　　(d) $Pt(NH_3)Cl_3$

20. Wilson's disease arises from:

    (a) Excess accumulation of calcium in the body

    (b) Excess accumulation of copper in the body

    (c) Excess accumulation of selenium in the body

    (d) Excess accumulation of vanadium in the body

21. The structure shown by a tetra nuclear metal cluster having 62 electrons is

    (a) square plane　　　　(b) butterfly　　　　(c) tetrahedron　　　　(d) bicapped tetrahedron

22. The substitution reactions of a square planar complex proceeds by which one of the following rate law:

    (a) first order　　　　(b) pseudo first order　　　　(c) second order　　　　(d) zero order

23. $[Zr(CH_3)_6]$ exists in

    (a) Octahedral geometry　　　　(b) Trigonal prismatic geometry

    (c) Square pyramidal geometry　　　　(d) Distorted trigonal bipyramidal geometry

24. Which one of the following pair shows variable valence?

    (a) Zr, Ti　　　　(b) Bi, In　　　　(c) Lu, Gd　　　　(d) Pd, Cd

25. Which one of the following plays a major role in EDTA complexometric titrations?

    (a) concentration of metal ion　　　　(b) concentration of ligand

    (c) nature of buffer　　　　(d) temperature of the reaction

26. Monochromatic X-rays having a wave length of 10.4Å are preferentially diffracted by a crystal at an angle 25.5°, assuming that this is the first order diffraction with a d-spacing between crystalline planes equal to

    12.1Å. What is the value of $\sin\theta$ for the angle for the second order diffraction?

    (a) 0.959　　　　(b) 0.759　　　　(c) 0.859　　　　(d) 0.659

27. Addition of bismuth chloride to excess of water produces

    (a) clear solution　　　　(b) yellow solution　　　　(c) white precipitate　　　　(d) orange red precipitate

28. Oxidation state of nitrogen is correctly given for

    (a) $\left[Co(NH_3)_5 Cl\right]Cl_2 \, (0)$
    (b) $NH_2OH \, (-1)$

    (c) $(N_2H_5)_2 SO_4 \, (-2)$
    (d) $Mg_3N_2 \, (+3)$

29. The equilibrium $Cr_2O_7^{2-} \rightleftharpoons 2CrO_4^{2-}$

    (a) exists in acidic medium
    (b) exists in basic medium
    (c) exists in neutral medium
    (d) does not exist

30. Of the following, which is acidic in nature?
    (a) MnO
    (b) Mn_2O_7
    (c) Mn_2O_3
    (d) MnO_2

31. Among pivaladehyde, furfural, formaldehyde and p-toluldehyde, the aldehydes that undergo Cannizaro reaction are
    (a) Formaldehyde only
    (b) p-toluldehyde
    (c) furfural and formaldehye
    (d) all four aldehyde

32. Product of the following reaction is
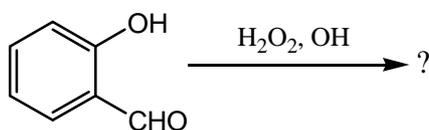
    (a) Catechol
    (b) Salicylic acid
    (c) o-benzoquinone
    (d) Salicyl alcohol

33. Which one of the following reagents are used as laboratory reagent for the identification of carbonyl group
    (a) NH_2OH/H^+
    (b) $C_6H_5\text{-}NH\text{–}NH_2/OH^-$

    (c)
    (d)
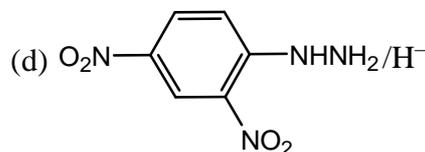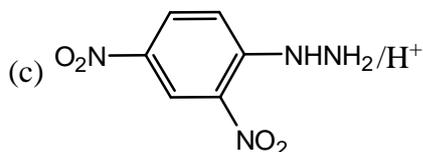
34. What is the major product alkene formed in the following reaction
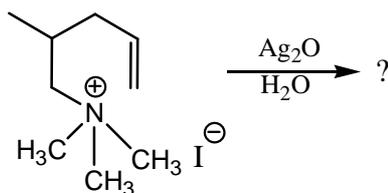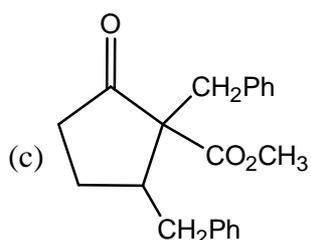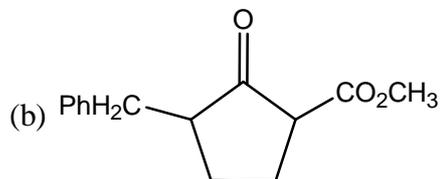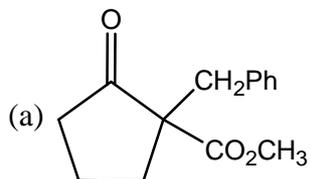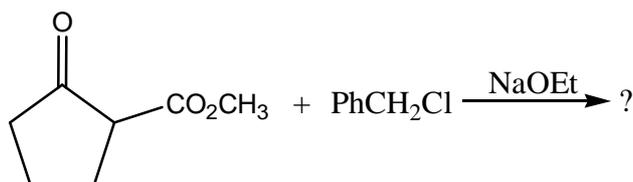

    (a)
    (b)
    (c)
    (d)

35. The suitable reagent(s) to bring about the given transformation
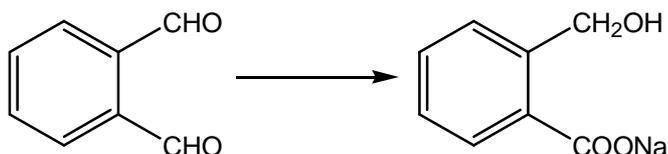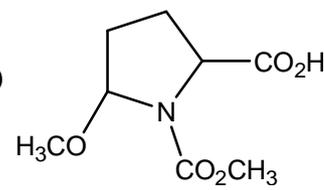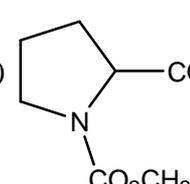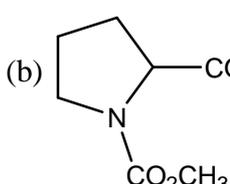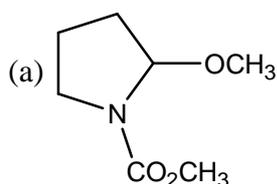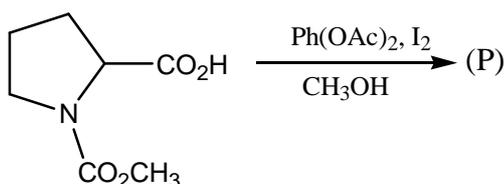

    (a) KCN
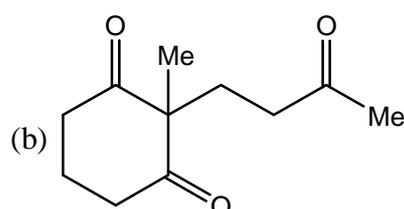   (b) $LiAlH_4$ followed by Tollens reagents
    (c) Tollens reagents followed by $NaBH_4$
    (d) 50% $NaOH$ solution

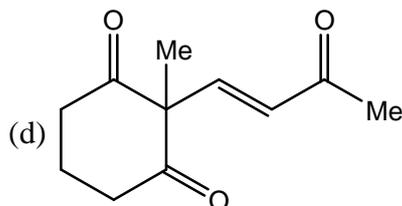36. Complete the following reaction,

$$\text{(cyclopentanone-2-CO}_2\text{CH}_3\text{)} + \text{PhCH}_2\text{Cl} \xrightarrow{\text{NaOEt}} ?$$

(a) 2-CH₂Ph, 2-CO₂CH₃ cyclopentanone structure
(b) PhH₂C and CO₂CH₃ substituted cyclopentanone

(c) CH₂Ph, CO₂CH₃ and CH₂Ph substituted cyclopentanone
(d) CO₂CH₃ and CH₂Ph substituted cyclopentanone
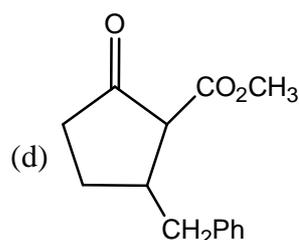
37. What is the product of this reaction?

$$\text{proline-N-CO}_2\text{CH}_3, \text{ -CO}_2\text{H} \xrightarrow[\text{CH}_3\text{OH}]{\text{Ph(OAc)}_2, \text{ I}_2} (P)$$

(a) pyrrolidine with -OCH₃ and N-CO₂CH₃
(b) pyrrolidine with -CO₂CH₃ and N-CO₂CH₃
(c) pyrrolidine with -COCH₃ and N-CO₂CH₃
(d) pyrrolidine with -CO₂H, H₃CO and N-CO₂CH₃

38. Which one is the correct answer?

$$\text{(2-Me-cyclohexane-1,3-dione)} + \text{(methyl vinyl ketone)} \xrightarrow[\text{1h, 75ºC}]{\text{AcOH/H}_2\text{O}} ?$$

(a) octahydronaphthalenedione with Me
(b) 2-Me-2-(3-oxobutyl)cyclohexane-1,3-dione

(c)

(d)
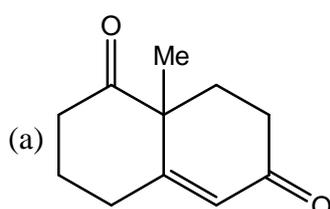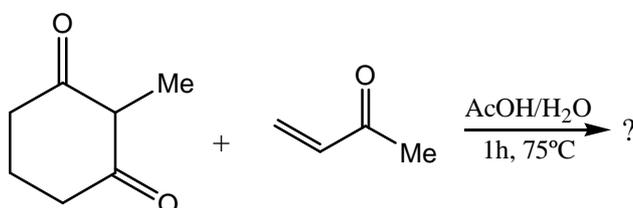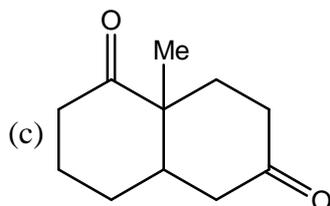
39. Match **List-(I)** with **List-(II)**

List-(I)	**List-(II)**
(P) Claisen condensation	(1) Nitrene
(Q) Friedal craft	(2) Free radical
(R) Holfmann	(3) Carbanion
(S) Allylic bromination	(4) Arenium ion

	P	Q	R	S
(a)	3	4	2	1
(b)	2	3	4	1
(c)	2	3	1	4
(d)	3	4	1	2

40. Match **List-(I)** with **List-(II)**

List-(I)	**List-(II)**
(P) S_N2	(1) Walden inversion
(Q) S_N1	(2) Carbanion intermediate
(R) E_2	(3) Antiperiplanar configurations
(S) E_{1CB}	(4) Carbocation intermediate

	P	Q	R	S
(a)	1	4	2	3
(b)	1	4	3	2
(c)	4	1	3	2
(d)	4	1	2	3

41. The configuration description of C_2 epimer of D-glucose is:

(a) 2R, 3S, 4R, 5R   (b) 2S, 3S, 4R, 5R

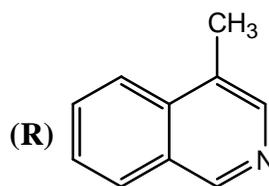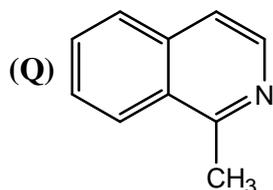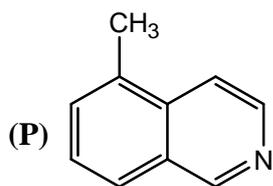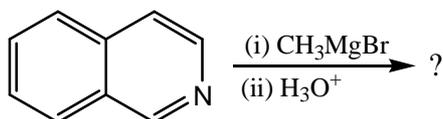(c) 2S, 3R, 4S, 5R   (d) 2R, 3S, 4R, 5S

42.
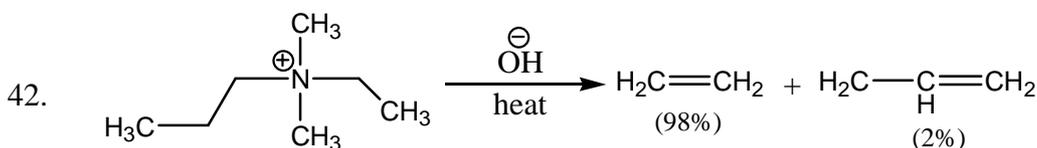
The above reaction is an example of

(a) Hofmann's rule   (b) Saytzeff's rule   (c) Cope reaction   (d) Curtius reaction
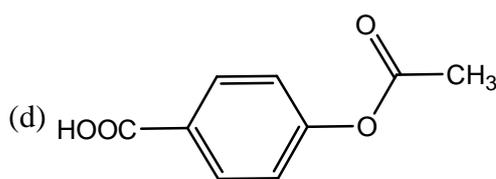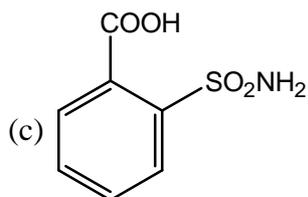
43. What is/are the product (s) of the following reaction?

(i) CH₃MgBr
(ii) H₃O⁺ → ?

(P)

(Q)

(R)

(a) P   (b) Q   (c) R   (d) P and Q

44. The structural formula of aspirin is

(a)

(b)

(c)

(d)
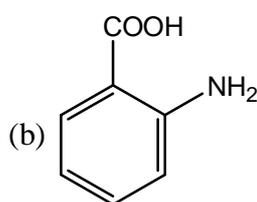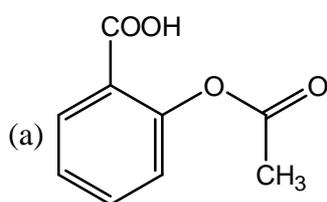
45. The compound that gives precipitate on warming with aqueous $AgNO_3$ is

(a)   (b)   (c)   (d)
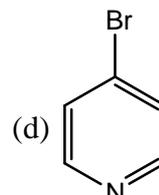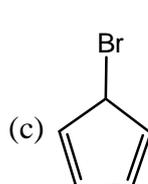
46. In the given reaction, name and the intermediate via which it is known to proceed are

(i) NaNO₂/HCl
(ii) CuCl/HCl
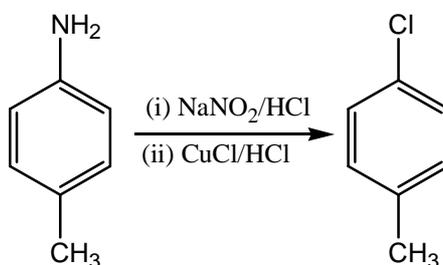
(a) Hunsdieecker and benzyne   (b) Sandmeyer and free radical

(c) Meerwein and a free radical   (d) Sandmeyer and carbanion

47. In the Cannizzaro reaction given below

$$HCHO + HCHO \xrightarrow{\text{NaOH}} CH_3OH + HCOONa$$

The slower step is

(a) the attack of OH– at carbonyl group   (b) The transfer of hydride ion to the carbonyl group
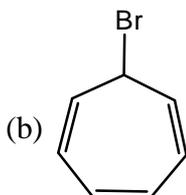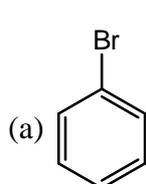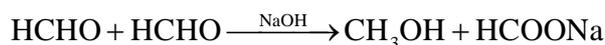
(c) Deprotonation of –COOH group   (d) The deprotonation of –CH_2OH group

48. In the given reaction, the structure of the σ complex is
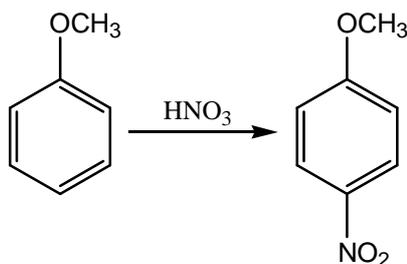
(a)   (b)   (c)   (d) all of these

49. How many non-equivalent protons are present in $CH_3CHClCH_2CONH_2$?

(a) 6   (b) 5   (c) 4   (d) 3

50. To check that a secondary alcohol has been completely oxidized to a ketone you can

(a) check that the IR spectrum has absorption at 3500 cm^{-1} and 1650 cm^{-1}

(b) check that the IR spectrum has no absorption around 3500 cm^{-1}

(c) check that the IR spectrum has no absorption around 1650 cm^{-1}

(d) check that the IR spectrum no absorptions at 3500 cm^{-1} and 1650 cm^{-1}

51. The major product formed in the reaction given below is
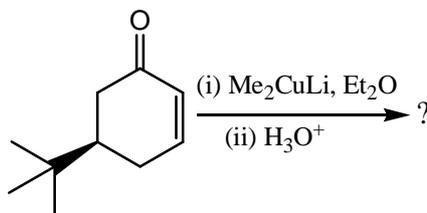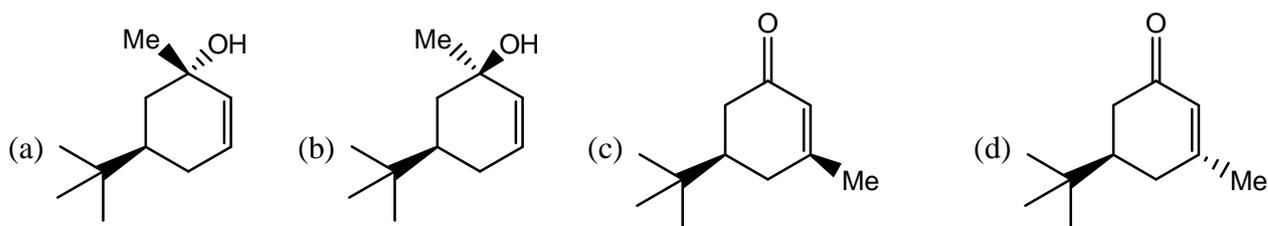
(a) [structure: Me, OH cyclohexene with isopropyl] (b) [structure: Me, OH cyclohexene with isopropyl] (c) [structure: ketone with Me] (d) [structure: ketone with Me]

52. The major product formed in the reaction given below is

[structure with NH₂] $\xrightarrow[\text{0–5ºC}]{\text{NaNO}_2,\ \text{aq. HCl}}$

(a) [structure with NO₂] (b) [structure with OH]

(c) [structure with NO₂] (d) [structure with OH]

53. In the reaction, Ph [structure with O and CHO] $\xrightarrow{\text{Ph}_3\text{P=CHCOOEt}}$ [X]

the major product X is:

(a) [structure H, COOEt, Ph, CHO] (b) [structure EtOOC, H, Ph, CHO]

(c) Ph [structure O, H, COOEt, H] (d) Ph [structure O, H, H, COOEt]
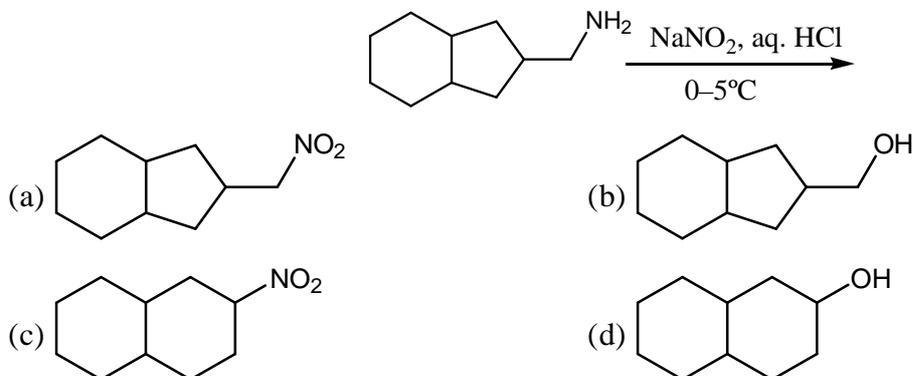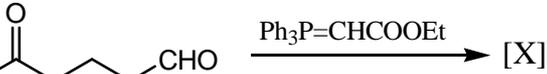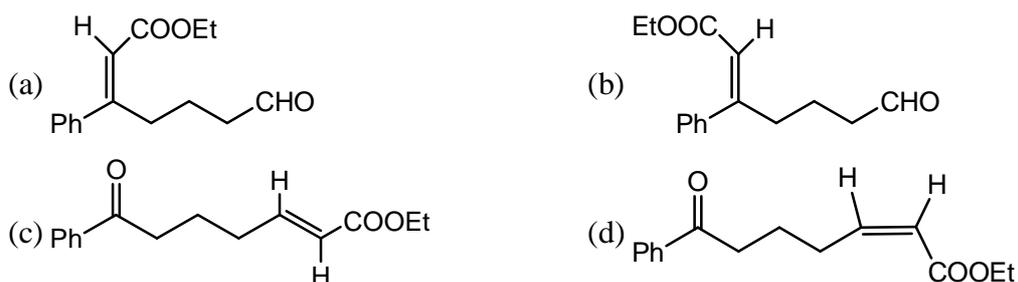
54. The major product (X) of the monobromination reaction is

Me [cyclohexane] $\xrightarrow[hv]{\text{Br}_2}$ (X)

(a) [Br-CH₂-cyclohexane] (b) Me [cyclohexane] Br (c) Me [Br cyclohexane] (d) [Br, Me cyclohexane]
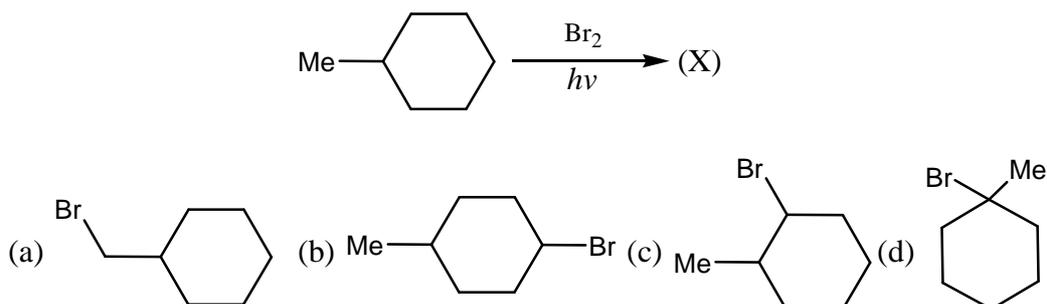
55. Aniline can be distinguished from methylamine by its reaction with :

(a) p-toluene sulphonyl chloride/KOH   (b) (i) NaNO₂/HCl (ii) alkaline β-naphthol

(c) Sn/HCl   (d) acetyl chloride

56. The most acidic species is

(a) (b) (c) (d)
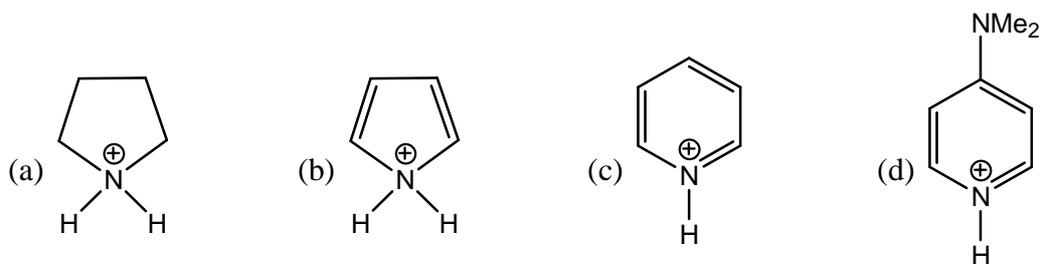
57. The major product of the following reactions is

$$\xrightarrow[\text{(ii) NaBH}_4, \text{OH}^-, \text{heat}]{\text{(i) Hg(OAc)}_2, \text{THF}} \quad ?$$

(a)

(b)

(c)

(d)
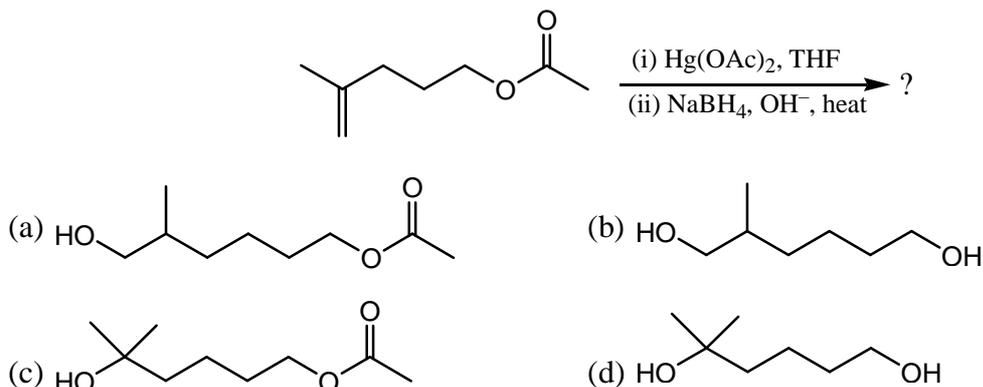
58. In the following reaction,

$$+ \quad Ph-CHO \quad \xrightarrow[\text{DMF}]{\text{NaCN(cat.)}} \quad [X]$$

The major product (X) is

(a) (b) (c) (d)
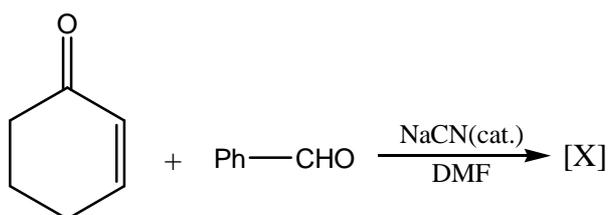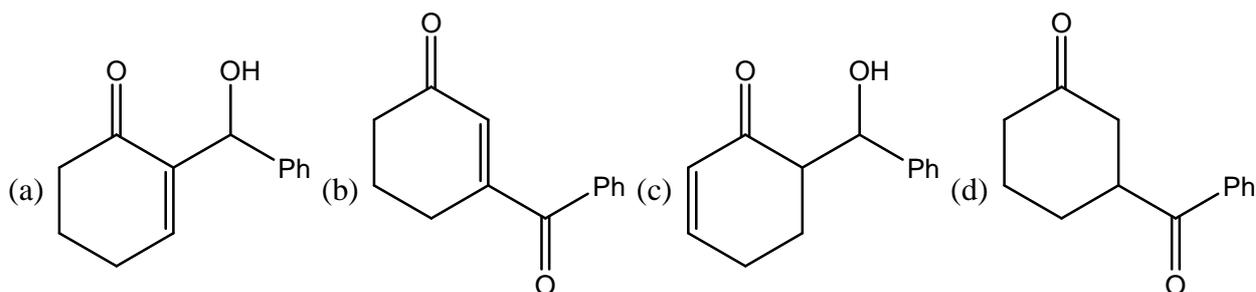
59. Correct statement about cis and transtilbene is

   (a) Trans-stibene has higher coupling constant than cis-stilbene

   (b) Cis-stibene has higher coupling constant than trans-stilbene

   (c) Vinylic protons of cis-stilbene are more deshielded than trans-stilbene

   (d) Vinylic protons of trans-stilbene are more deshielded than cis-stilbene

60. Victor Meyer test is used for the confirmation of

   (a) 1°, 2°, 3°, Amines   (b) 1°, 2°, 3° Alcohols

   (c) Carbonyl group   (d) 2° and 3° Alcohols only

61. The major product formed in the dinitration of 4-bromotoluene is

(a) [structure: toluene with NO2 at ortho (2) position and NO2 at para (4) position]

(b) [structure: toluene with two NO2 groups adjacent and Br]

(c) [structure: toluene ring with O2N and NO2 groups and Br]

(d) [structure: toluene with NO2, O2N and Br]
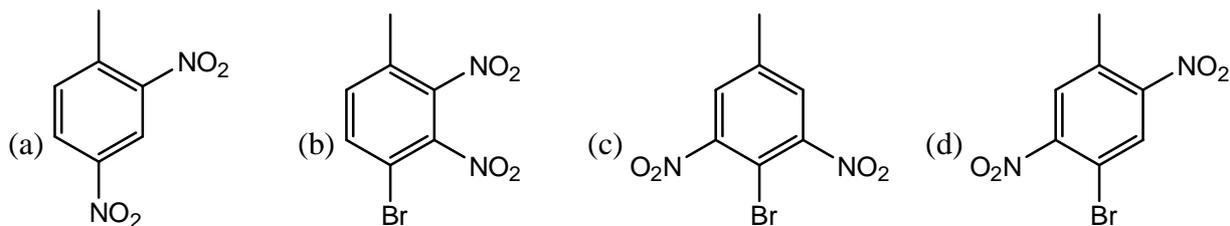
62. The major product in the following reaction is

[structure: (CH3)2C(Br)–C(=O)–CH2Br] $\xrightarrow[\text{Et}_2\text{O}]{\text{NaOMe}}$?

(a) [structure: H2C=C(Me)–C(=O)–CH2Br]

(b) [structure: cyclopropanone with two Me groups]

(c) [structure: Me2C=C(Me)–CO2Me]

(d) [structure: MeO2C–C(Me)2–CH2Br]
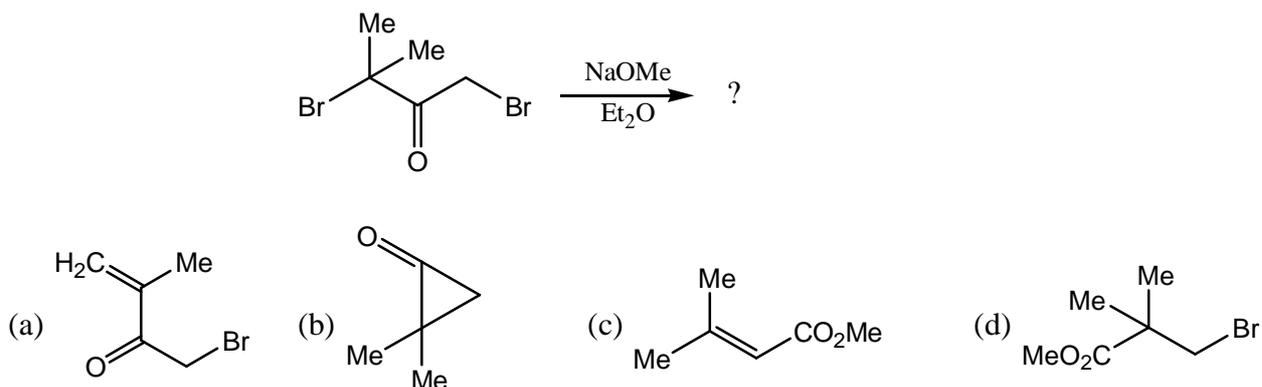
63. The decreasing order of basicity of the following compounds is
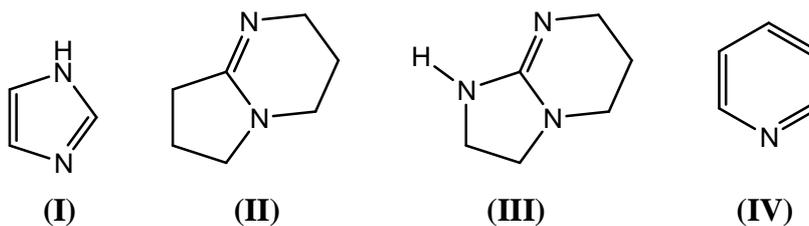
(I) **(II)** **(III)** **(IV)**

(a) I > II > III > IV

(b) IV > I > II > III

(c) III > II > I > IV

(d) IV > III > II > I

64. The correct order of pKa values for compounds I, II and III is
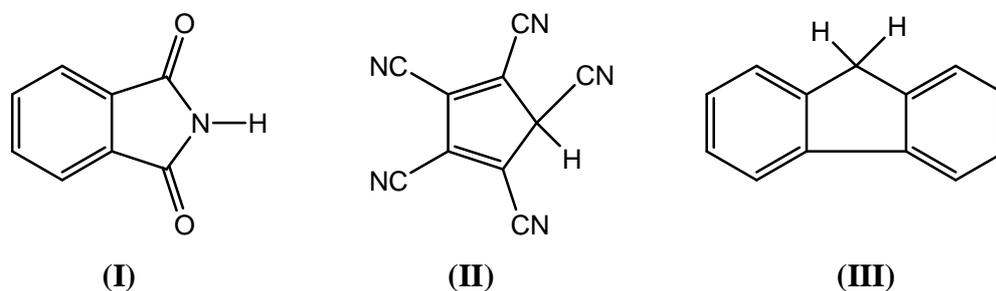
(I) **(II)** **(III)**

(a) I > II > III

(b) II > III > I

(c) III > I > II

(d) II > I > III

65. The correct order of the acidity for the following compounds is

        (I)   (II)   (III)

  (a) II > III > I   (b) III > II > I   (c) II > I > III   (d) III > I > II

66. The H-hydrogen elimination will be facile in :

  (a) 2   (b) 1   (c) 3   (d) 4

67. What is the degeneracy and corresponding energy for the second excited level of a particle of mass m in a cubic box of edge L?
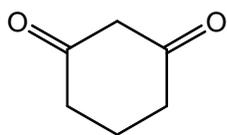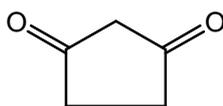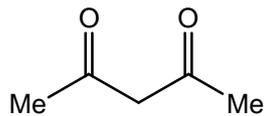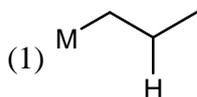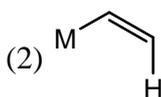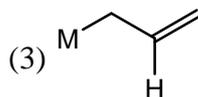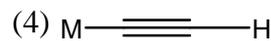
  (a) Degeneracy = 3 and $E = 9\pi^2 h^2 / 2mL^2$ (b) Degeneracy = 3 and $E = 4\pi^2 h^2 / 2mL^2$

  (c) Degeneracy = 2 and $E = 4\pi^2 h^2 / 2mL^2$ (d) Degeneracy = 2 and $E = 9\pi^2 h^2 / 2mL^2$

68. Which of the following statements is NOT true in relation to the triple point on a single component phase diagram?

  (a) The point at which the solid, liquid and gaseous phases for a substance co-exist

  (b) The triple point exists at a single temperature and is independent of pressure

  (c) The triple point exists for a substance occurs at a specific temperature and pressure

  (d) The system must be enclosed so that no vapour can escape

69. The relative ratio of .... at a given temperature is

  (a) $\sqrt{3} : \sqrt{56/22} : \sqrt{2}$ (b) $\sqrt{56/22} : \sqrt{2} : \sqrt{3}$

  (c) $\sqrt{3} : \sqrt{2} : \sqrt{56/22}$ (d) $\sqrt{2} : \sqrt{56/22} : \sqrt{3}$

70. The standard potential at 25ºC for the half reactions given against them below

$$Zn \rightarrow Zn^{2+} + 2e^- \qquad E^0 = 0.762$$

$$Mg \rightarrow Mg^{2+} + 2e^- \qquad E^0 = 2.37$$
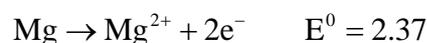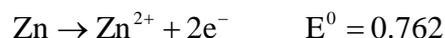
When zinc dust is added to a solution of $MgCl_2$

  (a) Magnesium is precipitated   (b) Zinc dissolves in the solution

  (c) Zinc chloride is formed   (d) No reaction takes place

71. Rate constant for the substitution reaction,

$$C_4H_9Cl + H_2O \longrightarrow C_4H_9OH + HCl$$

increases by a factor of 10.6 when the temperature is increased from 298K to 308K. Calculate the activation energy of the reaction,

  (a) 78.2 kJ mol^{-1} (b) 180 kJ mol^{-1} (c) 809 kJ mol^{-1} (d) 2.14 kJ mol^{-1}

72. At 20ºC, Ag ion concentration in a saturated solution Ag_2CrO_4 in water is 1.5×10^{-4} M. At 20ºC, the

solubility product of Ag_2CrO_4 will be
(a) 3.3750×10^{-12} (b) 1.6875×10^{-16} (c) 1.6875×10^{-12} (d) 1.6875×10^{-12}

73. Use the following data to calculate the lattice enthalpy at 298K of potassium iodide, KI(s). All values refer to a temperature of 298K.

Enthalpy of sublimation of K(s) : + 81 kJ mol^{-1}

Ionization of enthalpy of K(g) : +418 kJ mol^{-1}

Enthalpy of atomization of I_2(g) : +214 kJ mol^{-1}

Enthalpy of electron attachment to I(g) : –295 kJ mol^{-1}

Enthalpy of formation of KI(s) from K(s) and $\frac{1}{2}I_2$(g) : –328 kJ mol–1

(a) 746 kJ mol^{-1} (b) 680 kJ mol^{-1} (c) 573 kJ mol^{-1} (d) 639 kJ mol^{-1}

74. Which of the following functions are acceptable as a wave function over the indicated intervals?

(a) $e^{-|x|}\left(-\infty, \infty\right)$ (b) $e^{-x^2}\left(-\infty, \infty\right)$ (c) $e^{-x}\left(0, \infty\right)$ (d) $\sin^{-1} x\left(-1, 1\right)$

75. The thermodynamics Gibb's function G in an isothermal, isobaric, reversible process:
(a) varies linearly   (b) varies non-linearly   (c) is zero   (d) remains constant but not zero

76. The electrical conductivity of a new material was measured at different temperatures and found to vary as below. What is the best description of the conduction properties of the material?

T/K	300	400	500	600
Conductivity/Sm^{-1}	0.004	0.047	0.202	0.535

(a) insulator   (b) semiconductor   (c) conductor
(d) it is not possible to inter anything about the properties of the material

77. What miller index plane is shown below?
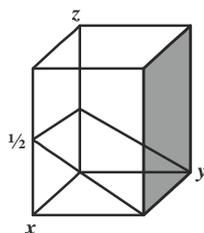
(a) (0 2 1)   (b) (0 4 1)   (c) (0 1 2)   (d) (0 0 0)

78. If uncertainties in the measurement of position and momentum are equal, calculate uncertainty in the measurement of velocity

(a) 6.96×10^8 m/s   (b) 7.98×10^{12} m/s   (c) 7.98×10^{14} m/s   (d) 7.98×10^{-12} m/s s

79. Following are the terms about activity and selectivity:
(I) Activity is the ability of catalyst to accelerate chemical reaction and selectivity is the ability of the catalysts to direct to reaction to yield particular products
(II) Activity is the ability of catalyst to direct reaction to yield particular products and selectivity is the ability of the catalysts to accelerate chemical reaction.
Select correct term :
(a) I   (b) II   (c) I and II both   (d) None of these

80. Which statement is incorrect about CO_2?
(a) CO_2 is linear

(b) CO_2 has two degenerate bending modes of vibration

(c) The IR spectrum of CO_2 shows four absorptions

(d) CO_2 has two stretching modes of vibration

81. From the Carnot cycle (given below) undergoes by an ideal gas, identify the processes in which the change in internal energy is non-zero
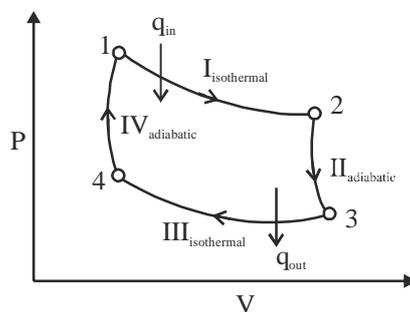
   (a) II and III   (b) I and II   (c) II and IV   (d) I and IV

82. For a liquid, which is rising in a capillary tube, the angle of contact is:

   (a) 90º   (b) Acute   (c) 180º   (d) Obtuse

83. Assign the Bravais lattice type for the following unit-cell structure
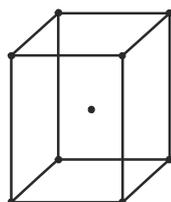
   (a) Tetragonal I   (b) Cubic I   (c) Orthorhombic I   (d) Monoclinic

84. Given the following reaction at equilibrium $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$. Some inert gas is added at constant volume. Predict which of the following facts will be affected?

   (a) Less of $2NH_3(g)$ is produced

   (b) No affect on the degree of advancement of the reaction at equilibrium

   (c) More of the $2NH_3(g)$ is produced

   (d) K_P of the reaction is increased.

85. On what factor fluidity of a liquid is depending

   (a) size of molecules   (b) Shape of the molecules

   (c) impurities   (d) all of these

86. Which of the following statements is correct about the principal moments of inertia of an XY molecules that lies on the A-axis?

   (a) $I_A = 0$ and $I_B = I_C$ (b) $I_A > I_B$ and $I_B = I_C$

   (c) $I_A > I_B$ and $I_C = 0$ (d) $I_A = I_B = I_C$

87. Two van der Waals gases have the same value of b but different a values. (i) Which of these would occupy greater volume under identical conditions? If the gases have the same a value but different values of b, (ii)

which would be more compressible?
(a) (i) gas with smaller a value (ii) gas with smaller b vlaue
(b) (i) gas with smaller a value (ii) gas with larger b value
(c) (i) gas with smaller a value (ii) gas with smaller b value
(d) (i) gas with larger a value (ii) gas with larger b value

88. Consider the following Potential energy diagram for a reversible reaction
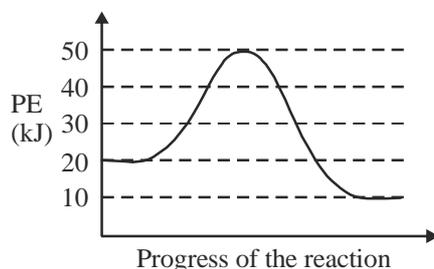
Which of the following describes this reaction:

	Directions	Activation energy (kJ)	$\Delta H(kJ)$
(a)	Reverse	30	−10
(b)	Forward	40	−10
(c)	Forward	30	+10
(d)	Reverse	40	+10

89. Which of the following statements are true about the Eutectic point on a two component (components A and B) phase diagram?
(a) Both compounds are solid
(b) The melting point of the mixture is lower than the melting points of either of the individual compounds
(c) One compound is in the liquid phase while the other is in the solid phase
(d) It always occurs when the ratio of compound A to compound B is 50.50

90. The reference potential of a silver-silver chloride electrode is determined by the
(a) Concentration of potassium chloride filling solution
(b) Surface area of the electrode
(c) Activity of total anion in the paste covering electrode
(d) Concentration of silver in the paste covering electrode

91. Which statement is not correct about three equilibrium constants, K_c, K_p and K_x

(a) $K_p = K_x (P)^{\Delta n} = K_c (RT)^{\Delta n}$

(b) K_c and K_p are independent of pressure and K_x is dependent on pressure.
(c) All constants are temperature dependent
(d) Catalysts change the equilibrium

92. pH of the solution produced by mixing equal volumes of 2.0×10^{-3} M $HClO_4$ and 1.0×10^{-3} M $KClO_4$ is :
(a) 3.0 　　　(b) 2.7 　　　(c) 2.3 　　　(d) 1.0

93. The degree of dissociation (α) of a weak electrolyte, $A_x B_x$ is related to van't Hof factor (s) by the expressionns:

(a) $\alpha = \dfrac{i-1}{(x+y+1)}$ (b) $\alpha = \dfrac{i-1}{(x+y-1)}$ (c) $\alpha = \dfrac{x+y-1}{i-1}$ (d) $\alpha = \dfrac{x+y+1}{i-1}$

94. For a reaction involving two steps given below

First step : $G \rightleftharpoons 2H$

Second step: $G + H \longrightarrow P$

Assume that the first step attains equilibrium rapidly. The rate of formation of P is proportional to

(a) $[G]^{1/2}$ (b) $[G]^{2}$ (c) $[G]^{3/2}$ (d) $[G]$

95. 1g of ^{86}Sr gets converted to 0.953g after 2 year. The half life of ^{90}Sr and the amount of ^{85}Sr remainingafter 5 years are
(a) 28.8 yr and 0.887 g (b) 1.44 yr and 0.75g
(c) 57.6 yr and 0.75 g (d) 100 yr and 0.982 g

96. The normality of 0.3 M phosphorons acid (H_3PO_3) is
(a) 0.9 (b) 0.6 (c) 0.3 (d) 0.1

97. Lattice energy is decreased when size of anion is
(a) decreased (b) increased (c) remains same (d) no change

98. One of angle in monoclinic crystal system is
(a) less than 90 degree (b) greater than 90 degree
(c) less than 30 degree (d) less than 10 degree

99. An oxidation number can be
(a) positive (b) negative (c) zero (d) all of the above

100. Liquid in an electrolyte cell should always flow towards
(a) anode (b) cathode (c) all around (d) nowhere